Name:

## Understanding, Making, and Using Buffers: Part II: Practicing Titrations

## **DATA SHEET**

| Beaker 1 with 50-mL water:  | Record pH of water  | Record pH upon addition of 1-mL HCl  |
|-----------------------------|---------------------|--------------------------------------|
| Beaker 2 with 50-mL water:  | Record pH of water  | Record pH upon addition of 1-mL NaOH |
| Beaker 1 with 50-mL buffer: | Record pH of buffer | Record pH upon addition of 1-mL HCl  |
| Beaker 2 with 50-mL buffer: | Record pH of buffer | Record pH upon addition of 1-mL NaOH |

1. What did your group observe in this part of the lab?

2. How does a buffer solution resist a change in pH?

3. Why would HCl and NaOH be a poor choice for an acid-base pair to make a buffer?